

Subject card

Subject name and code	Inorganic chemistry, PG_00080829						
Field of study	Chemical Business						
Date of commencement of studies	October 2026		Academic year of realisation of subject		2026/2027		
Education level	Bachelor's studies		Subject group		Obligatory subject group in the field of study Subject group related to scientific research in the field of study		
Mode of study	full-time studies		Mode of delivery		at the university		
Year of study	1		Language of instruction		Polish		
Semester of study	2		ECTS credits		2.0		
Learning profile	academic		Assessment form		exam		
Conducting unit	Laboratory of Physicochemistry of Coordination Complexes -> Department of General and Inorganic Chemistry -> Faculty of Chemistry -> Rector						
Name and surname of lecturer (lecturers)	Subject supervisor		dr hab. Dariusz Wyrzykowski				
	Teachers						
Lesson types	Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
	Number of study hours	30.0	0.0	0.0	0.0	0.0	30
	E-learning hours included: 0.0						
Learning activity and number of study hours	Learning activity	Participation in didactic classes included in study plan		Participation in consultation hours		Self-study	SUM
	Number of study hours	30		5.0		15.0	50
Subject objectives	- to consolidate basic theoretical knowledge of inorganic chemistry - to introduce important problems of modern inorganic chemistry - introduction to the most important contemporary issues in inorganic chemistry which constitute progress in this field - to develop the ability to conduct experiments independently, to interpret the results obtained and to solve problems while conducting chemical experiments						
Learning outcomes	Course outcome		Subject outcome		Method of verification		
	[BCHINŻ_U08] Uses the chemical nomenclature and engineering terminology properly.		Appropriately uses chemical nomenclature and engineering terminology.		[SU4] test/exam - oral or written		
	[BCHINŻ_U02] Uses basic methods, techniques and tools in formulating and solving engineering tasks in the field of chemistry.		Applies methods, techniques and tools in formulating and solving engineering tasks in the field of chemistry.		[SU4] test/exam - oral or written		
	[BCHINŻ_U09] Using the acquired knowledge, skills and various sources of scientific information independently prepares written papers and oral presentations.		Using the acquired knowledge, skills and various sources of scientific information, he independently prepares written works and oral presentations.		[SU4] test/exam - oral or written		
	[BCHINŻ_W02] Enumerates basic laws and theories in chemistry, physics and mathematics necessary to formulate and solve simple engineering tasks.		Enumerates basic laws and theories in chemistry, physics and mathematics necessary to formulate and solve simple engineering tasks,		[SW3] text preparation/written work		

Subject contents	Topics of the lecture: periodicity and the chemistry of the elements, physicochemical properties of inorganic and coordination compounds. The following items are included: periodicity, chemical bonding, coordination compounds, types of chemical reactions, properties of chemical elements and their compounds. The groups of elements are presented in the following order: group 1, group 2, group 13, group 14, group 15, group 16, group 17, group 18, and d-elements (groups 3-12; first transition row, second transition row, and third transition row). Topics of auditory classes: basic types of inorganic compounds, valence bond theory, hybridization and molecular geometry; molecular orbital theory; solid state bonds: ionic, covalent, metallic; metals, semiconductors and insulator; coordination compounds. Topics of lab classes: investigation of physicochemical properties of the elements, inorganic and coordination compounds based on chemical experiments.		
Prerequisites and co-requisites			
Assessment methods and criteria	Subject passing criteria	Passing threshold	Percentage of the final grade
		51.0%	100.0%
Recommended reading	Basic literature	A. Bielański Podstawy chemii nieorganicznej, PWN 2002 L. Jones, P. Atkins Chemia ogólna, PWN 2004 B. Literatura uzupełniająca L. Pajdowski Chemia ogólna, PWN 1999	
	Supplementary literature	J. D. Lee Związła chemia nieorganiczna, PWN 1997	
	eResources addresses		
Example issues/ example questions/ tasks being completed	Based on the theory of hard and soft acids and bases, explain the difference in solubility of silver halides. Explain the differences in magnetic properties (diamagnetic, paramagnetic) of sodium peroxide and potassium superoxide (Hint: based on an analysis of the electron distribution on the molecular orbitals of the peroxide ion and the superoxide ion, indicate which of the individuals has unpaired electrons/electrons. On this basis, indicate which of the compounds exhibits paramagnetic and which exhibits diamagnetic properties).		
Work placement	Not applicable		

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