

Subject card

Subject name and code	General chemistry, PG_00053406						
Field of study	Chemistry						
Date of commencement of studies	October 2026		Academic year of realisation of subject		2026/2027		
Education level	Bachelor's studies		Subject group		Obligatory subject group in the field of study Subject group related to scientific research in the field of study		
Mode of study	full-time studies		Mode of delivery		at the university		
Year of study	1		Language of instruction		Polish		
Semester of study	1		ECTS credits		4.0		
Learning profile	academic		Assessment form		exam		
Conducting unit	Laboratory of Physicochemistry of Coordination Complexes -> Department of General and Inorganic Chemistry -> Faculty of Chemistry -> Rector						
Name and surname of lecturer (lecturers)	Subject supervisor		dr hab. Joanna Makowska				
	Teachers						
Lesson types	Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
	Number of study hours	45.0	0.0	0.0	0.0	0.0	45
	E-learning hours included: 0.0						
Learning activity and number of study hours	Learning activity	Participation in didactic classes included in study plan		Participation in consultation hours		Self-study	SUM
	Number of study hours	45		8.0		47.0	100
Subject objectives	- to consolidate basic theoretical knowledge of inorganic chemistry - to introduce important problems of modern inorganic chemistry - introduction to the most important contemporary issues in inorganic chemistry which constitute progress in this field - to develop the ability to conduct experiments independently, to interpret the results obtained and to solve problems while conducting chemical experiments						

Learning outcomes	Course outcome	Subject outcome	Method of verification
	[CHEML3_W01] Enumerates basic laws and theories in chemistry, physics, mathematics and biology.	Lists laws and theories in chemistry, physics, mathematics and biology.	[SW4] test/exam - oral or written
	[CHEML3_U01] Identifies, analyses and solves problems in the field of broadly understood chemistry on the basis of the acquired knowledge.	Identifies, analyzes and solves problems in the field of broadly understood chemistry based on acquired knowledge.	[SU4] test/exam - oral or written
	[CHEML3_W08] Demonstrates knowledge of basic computational methods to solve problems in chemistry, physics, mathematics.	Demonstrates knowledge of computational methods for solving problems in chemistry, physics and mathematics.	[SW4] test/exam - oral or written
	[CHEML3_U08] Presents in an understandable way the basic facts about chemistry using a scientific language typical of chemical sciences.	It presents chemistry facts in an accessible way, using scientific language typical of chemical sciences.	[SU4] test/exam - oral or written
	[CHEML3_W13] Enumerates and describes the basic legal and ethical aspects related to scientific, research and didactic work.	Lists and describes the basic legal and ethical aspects related to scientific, research and teaching work.	[SW4] test/exam - oral or written
	[CHEML3_W14] Recalls and explains the basic concepts and principles in the field of intellectual and industrial property protection, copyright and patent law.	Summons and explains basic concepts and principles in the field of intellectual and industrial property protection, copyright and patent law.	[SW4] test/exam - oral or written
Subject contents	Topics of the lecture: atomistic theory of matter (atomic nucleus, isotopes, electronic structure of atoms, quantum numbers, atomic orbitals), basic chemical terms and laws, periodic table of elements, chemical equations (including redox reactions), chemical bonds, basic types of inorganic compounds, stoichiometry, solutions and their concentrations, thermochemistry, kinetics and chemical equilibrium, theories of acids and bases, electrolytic dissociation, pH scale, pH of solutions of strong and weak acids and bases, buffer solutions, hydrolysis, elements of electrochemistry. Topics of auditory classes: basic chemical terms and laws, basic types of inorganic compounds, balancing redox reactions, stoichiometry, the concentrations of the solutions, kinetics and chemical equilibrium, equilibria in the solutions of electrolytes.		
Prerequisites and co-requisites			
Assessment methods and criteria	Subject passing criteria	Passing threshold	Percentage of the final grade
		51.0%	100.0%
Recommended reading	Basic literature	Praca zbiorowa - Obliczenia z chemii ogólnej - skrypt UG, Gdańsk 2011	
		Praca zbiorowa - Ćwiczenia laboratoryjne z chemii ogólnej. I Część teoretyczna	
		Praca zbiorowa - Ćwiczenia laboratoryjne z chemii ogólnej. II Część doświadczalna	
	Supplementary literature	A. Bielański Podstawy chemii nieorganicznej	
		J. D. Lee Związła chemia nieorganiczna, PWN 1997	
		L. Jones, P. Atkins Chemia ogólna, PWN 2004	
	eResources addresses		
Example issues/ example questions/ tasks being completed	brak		
Work placement	Not applicable		

Document generated electronically. Does not require a seal or signature.