

Subject card

Subject name and code	General chemistry, PG_00053416						
Field of study	Chemistry						
Date of commencement of studies	October 2026		Academic year of realisation of subject		2026/2027		
Education level	Bachelor's studies		Subject group		Obligatory subject group in the field of study Subject group related to scientific research in the field of study		
Mode of study	full-time studies		Mode of delivery		at the university		
Year of study	1		Language of instruction		Polish		
Semester of study	1		ECTS credits		2.0		
Learning profile	academic		Assessment form		credit		
Conducting unit	Laboratory of Physicochemistry of Coordination Complexes -> Department of General and Inorganic Chemistry -> Faculty of Chemistry -> Rector						
Name and surname of lecturer (lecturers)	Subject supervisor		dr hab. Joanna Makowska				
	Teachers						
Lesson types	Lesson type	Lecture	Tutorial	Laboratory	Project	Seminar	SUM
	Number of study hours	0.0	0.0	30.0	0.0	0.0	30
	E-learning hours included: 0.0						
Learning activity and number of study hours	Learning activity	Participation in didactic classes included in study plan		Participation in consultation hours		Self-study	SUM
	Number of study hours	30		5.0		15.0	50
Subject objectives	- to consolidate basic theoretical knowledge of inorganic chemistry - to introduce important problems of modern inorganic chemistry - introduction to the most important contemporary issues in inorganic chemistry which constitute progress in this field - to develop the ability to conduct experiments independently, to interpret the results obtained and to solve problems while conducting chemical experiments						

Learning outcomes	Course outcome	Subject outcome	Method of verification
	[CHEML3_W01] Enumerates basic laws and theories in chemistry, physics, mathematics and biology.	Lists laws and theories in chemistry, physics, mathematics and biology.	[SW4] test/exam - oral or written
	[CHEML3_U08] Presents in an understandable way the basic facts about chemistry using a scientific language typical of chemical sciences.	It presents chemistry facts in an accessible way, using scientific language typical of chemical sciences.	[SU1] oral statement/conversation/discussion
	[CHEML3_W08] Demonstrates knowledge of basic computational methods to solve problems in chemistry, physics, mathematics.	Demonstrates knowledge of computational methods for solving problems in chemistry, physics and mathematics.	[SW4] test/exam - oral or written
	[CHEML3_U01] Identifies, analyses and solves problems in the field of broadly understood chemistry on the basis of the acquired knowledge.	Identifies, analyzes and solves problems in the field of broadly understood chemistry based on acquired knowledge.	[SU6] demonstration of practical skills
	[CHEML3_W12] Characterises the basic principles of health and safety at work in a chemical laboratory; knows and describes the hazards associated with working with hazardous substances, ways to counteract these hazards and rules of conduct during an accident.	Characterizes the principles of occupational health and safety in a chemical laboratory; knows and describes the threats related to working with hazardous substances, methods of counteracting these threats and rules of conduct in the event of an accident.	[SW5] implementation of a problem task
	[CHEML3_K03] Establishes priorities in the right way for the implementation of tasks specified by herself/himself and/or by others.	Sets appropriate priorities for the implementation of tasks specified by himself and/or others.	[SK8] observation of student's independent or team work
	[CHEML3_U03] Selects the appropriate equipment and laboratory apparatus for conducting uncomplicated chemical experiments.	Selects appropriate equipment and laboratory equipment for conducting chemical experiments.	[SU8] observation of student's independent or team work
Subject contents	Topics of the lecture: atomistic theory of matter (atomic nucleus, isotopes, electronic structure of atoms, quantum numbers, atomic orbitals), basic chemical terms and laws, periodic table of elements, chemical equations (including redox reactions), chemical bonds, basic types of inorganic compounds, stoichiometry, solutions and their concentrations, thermochemistry, kinetics and chemical equilibrium, theories of acids and bases, electrolytic dissociation, pH scale, pH of solutions of strong and weak acids and bases, buffer solutions, hydrolysis, elements of electrochemistry. Topics of auditory classes: basic chemical terms and laws, basic types of inorganic compounds, balancing redox reactions, stoichiometry, the concentrations of the solutions, kinetics and chemical equilibrium, equilibria in the solutions of electrolytes.		
Prerequisites and co-requisites			
Assessment methods and criteria	Subject passing criteria	Passing threshold	Percentage of the final grade
		51.0%	100.0%
Recommended reading	Basic literature Praca zbiorowa - Obliczenia z chemii ogólnej - skrypt UG, Gdańsk 2011 Praca zbiorowa - Ćwiczenia laboratoryjne z chemii ogólnej. I Część teoretyczna Praca zbiorowa - Ćwiczenia laboratoryjne z chemii ogólnej. II Część doświadczalna		

	Supplementary literature	A. Bielański Podstawy chemii nieorganicznej J. D. Lee Związła chemia nieorganiczna, PWN 1997 L. Jones, P. Atkins Chemia ogólna, PWN 2004
	eResources addresses	
Example issues/ example questions/ tasks being completed	brak	
Work placement	Not applicable	

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